

**Distribution of Topics for Mid-Term, First-Term and Second-Term**  
**(From June 2016 Onwards)**

**Subject: CHEMISTRY**

<b>Sr.No.</b>	<b>Topic</b>	<b>Mid-Term</b>	<b>First-Term</b>	<b>Second-Term</b>
1.	Unit 1:Some Basic Concepts Of Chemistry	06	05	
2.	Unit 2:Structure Of Atom	08	07	
3.	Unit 3:Classification Of Elements And Periodicity of Properties	06	05	
4.	Unit 4:Chemical Bonding And Molecular Structure		10	
5.	Unit 5:States of Matter		09	
6.	Unit 6:Thermodynamics		08	
7.	Unit 12:Organic Chemistry-Some Basic Principles and Techniques		11	
8.	Unit 7:Equilibrium			07
9.	Unit 8:Redox Reactions			07
10.	Unit 9: Hydrogen			05
11.	Unit 10:s-Block Elements			05
12.	Unit 11: p-Block Elements			06
13.	Unit 13:Hydrocarbons			12
14.	Unit 14:Environmental Chemistry			02
15.	Core Content			11
	<b>Total</b>	<b>20</b>	<b>55</b>	<b>55</b>

<b>Weightage To Objectives</b>		
1	Knowledge	25%
2	Understanding and Skill	50%
3	Application	25%

Goa Board of Secondary & Higher Secondary Education

Alto, Betim – Goa

CHEMISTRY -XI SCIENCE (FROM JUNE 2016 ONWARDS)

First Terminal Practical Examination

MODEL QUESTION PAPER

Date:

Duration: 2 Hours

Session:

Max. Marks: 20

**Instructions:**

- 1) Write your Examination Seat number and laboratory table number on your answer book
- 2) Get the burette reading and confirmatory tests initialed by the examiner.
- 3) Check if the number on (i) your table (ii) answer script and (iii) the containers A, B, and E/F are the same. If not, report immediately to the examiner.
- 4) Use of non - programmable calculator is allowed.

**Atomic Masses:-** H=1, C=12, N=14, O=16, Na =23, S=32, K=39.

Q.1. You are provided with two solutions as follows:-

**Container A:** \_\_\_\_\_ N/M solution of \_\_\_\_\_.

**Container B:** \_\_\_\_\_ Solution

Using the above solutions , determine:

i)N/M of the solution in container **B**.

ii)Convert the N/M obtained above to M/N.

iii)Calculate the strength of the solution in container **B** in terms of **g per \_\_\_\_\_ mL.**

**14Mk**

Q.2. Identify the natural organic substance supplied to you in Container **E/F** bearing your table number. (

**2Marks)**

Q.3. Journal + Viva

**(2+2**

**Marks)**

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**CHEMISTRY -XI SCIENCE**

**Second Term Practical Examination(FROM JUNE 2016 ONWARDS)**

**MODEL QUESTION PAPER**

Date:

Duration: 3

Hours

Session:

Max. Marks:40

**Instructions:**

- 1) Write your Examination Seat number and laboratory table number on your answer book
- 2) Get the burette reading and confirmatory tests initialed by the examiner.
- 3) Check if the number on (i) your table (ii) answer script and (iii) the containers A, B, and D are the same. If not, report immediately to the examiner.
- 4) Use of non - programmable calculator is allowed.

**Atomic Masses:- H=1, C=12, N=14, O=16, Na =23, S=32, K=39.**

Q.1. You are provided with two solutions as follows:-

**Container A:** \_\_\_\_\_ g / \_\_\_\_\_ mL solution of \_\_\_\_\_.

**Container B:** Impure solution of \_\_\_\_\_

Using the above solutions ,determine:

i)N/M of the solution in container **B**.

ii)Calculate the percentage purity of the solution in container **B**, \_\_\_\_\_ g of which have been dissolved per \_\_\_\_\_ mL. **(05 Marks)**

Q.2. Analyse the inorganic salt given in container **D** bearing your table number qualitatively and detect the cation and anion present. Give a complete report of all the tests performed. Write the formula of the

Compound detected.

**(10**

**Marks)**

Q.3. Journal + Viva

**(3+2**

**Marks)**

Q.4. Project + Viva  
Marks)

(12+8

**CHEMISTRY -XI SCIENCE**

**First Terminal Practical Examination (FROM JUNE 2016 ONWARDS)**

**Marking Scheme**

**Q1 Volumetric Analysis ( Acid-base titration):**  
Marks)

(14

i) Observations : (Burette, pipette, indicator, color change) (1Mark)

ii) Chemical equation (1Mark)

iii) Reading: (7 Marks)

- |              |       |
|--------------|-------|
| 1. $\pm 0.1$ | 7 mks |
| 2. $\pm 0.2$ | 6 mks |
| 3. $\pm 0.3$ | 5 mks |
| 4. $\pm 0.4$ | 4 mks |
| 5. $\pm 0.5$ | 3 mks |
| 6. $\pm 0.6$ | 2 mks |

iv) Calculations: (5Marks)

1.Determination of N/M 2 mks

2.Conversion of M/N 1 mk

3.Strength 2 mks

**Q2. Natural Organic substance**

**(2 Marks)**

**Q3.Journal and viva**

**(2+2 Marks)**

## Second Term Practical Examination

**Q 1) Volumetric Analysis ( Acid-base titration):**  
**Marks)**

**(05**

i) Observations : (Burette, pipette, indicator, color change) (  $\frac{1}{2}$ Mark)

ii)Reading (3 Marks)

± 0.1	3mks
± 0.2	2 $\frac{1}{2}$ mks
± 0.3	2mks
± 0.4	1 $\frac{1}{2}$ mks
± 0.5	1mk

iv)Calculations: (1  $\frac{1}{2}$  Marks)

1. Conversion of g/ V mL to M/N of Solution A  $\frac{1}{2}$  mk
2. Determination of N/M of Solution B  $\frac{1}{2}$  mk
3. Percentage purity of Solution B  $\frac{1}{2}$  mk

**Q2. Inorganic Qualitative analysis**  
**Marks)**

**(10**

1) Preliminary test /Dry test  $\frac{1}{2} \times 8 = 4$  mks

2) Wet test for anion:

Correct anion	(1 $\frac{1}{2}$ mks)
C.T	(1mk)

3) Wet test for cation:

Correct group	(2mks)
Correct cation	( $\frac{1}{2}$ mk)
C.T	( $\frac{1}{2}$ mk)

4. Correct formula (1/2 mk)

**Q3. Journal and viva**  
**Marks)**

**(3+2**

**Q4. Project and viva**  
**Marks)**

**(12+8**